HSSC

CHEMISTRY

SECTION – A (Marks 17)

Time allowed: 25 Minutes

Section – A is compulsory. All parts of this section are to be answered on this page and handed over to the Centre Superintendent. Deleting/overwriting is not allowed.

Do not use lead pencil.

حصّہ اوّل لازمی ہے۔ اس کے جوابات اسی صفحہ پر دے کر ناظم مرکزکےحوالے کریں۔ کاٹ کر دوبارہ لکھنے کی اجازت نہیں ہے۔

لیڈ پنسل کا استعمال ممنوع ہے۔

ROLL NUMBER						

Version No.			
9	9	9	9

Answer Sheet No. _____

ہر سوال کے سامنے دیے گئے، کریکولم کے مطابق درست دائرہ کو پر کریں۔ **Invigilator Sign. _____**

Fill the relevant bubble against each question according to curriculum: **Candidate Sign. _____**

	Question	A	B	C	D	A	B	C	D
1.	Suppose, a salt bridge is **NOT** used between the two half cells of a galvanic cell then the voltage will:	Decrease very slowly	Not change	Drop to zero	Become negative	◯	◯	◯	◯
2.	If a magnesium ribbon is placed in aqueous solution of silver nitrate $(AgNO_3)$	MgO is precipitated out	NO_2 will form	No reaction will occur	Silver is precipitated out	◯	◯	◯	◯
3.	A substance that has a greater tendency to accept electrons, it will have a greater:	Reduction potential	Redox potential	EMF value	Oxidation potential	◯	◯	◯	◯
4.	The mass percent of magnesium in a formula unit of magnesium oxide (MgO) is:	30	40	60	20	◯	◯	◯	◯
5.	The relative atomic mass of chlorine is 35.5. The mass of 2 moles of chlorine gas will be:	$71g$	$35.5g$	$18.75g$	$142g$	◯	◯	◯	◯
6.	The mass of 1 mole of ammonia gas is $17g$. Calculate the mass of 0.3 mole of ammonia.	$1g$	$1.7g$	$5.1g$	$0.15g$	◯	◯	◯	◯
7.	The energy of 1st orbit in hydrogen atoms is:	$-2.189\times10^{-18}\,J$	$-3.18\times10^{-18}\,J$	$3.18\times10^{-12}\,J$	$1.18\times10^{-18}\,J$	◯	◯	◯	◯
8.	The electronic configuration of carbon atom in its **excited** state is:	$1s^2, 2s^1, 2p^3$	$1s^1, 2s^0, 2p^5$	$1s^2, 2s^1, 2p^4$	$1s^2, 2s^2, 2p^2$	◯	◯	◯	◯
9.	The principle and azimuthal quantum number values for $2p$ orbitals are:	$n=1, l=2$	$n=1, l=0$	$n=2, l=0$	$n=2, l=1$	◯	◯	◯	◯
10.	The gas which is most ideal among the following at STP is:	H_2S	H_2	NH_3	SO_2	◯	◯	◯	◯
11.	Which one of the following molecules has dipole moment?	BeF_2	H_2O	N_2	CH_4	◯	◯	◯	◯
12.	One of the following phenomena which involves the weakening of intermolecular attractive forces is:	Condensation	Crystallization	Evaporation	Freezing	◯	◯	◯	◯
13.	The coordination number of each Na^+ in $NaCl$ crystal is:	6	8	12	4	◯	◯	◯	◯
14.	The atomic ratio of both of the crystals, $NaCl$ and CaO, is $1:1$, such property is called:	Allotropy	Isomorphism	Polymorphism	Anisotropy	◯	◯	◯	◯
15.	In a reversible reaction at equilibrium, if the concentration of any of the reactant is doubled, then the equilibrium constant would be:	Halved	Doubled	Unchanged	One fourth	◯	◯	◯	◯
16.	When the air is bubbled through pure water, the pH is lowered from 7 to 5.6, which gas in the air is responsible for the change?	Carbon dioxide	Nitrogen	Oxygen	Argon	◯	◯	◯	◯
17.	If $2g$ of $NaOH$ is present in $200ml$ of a solution, the molarity of this solution will be:	0.5	1	2	0.25	◯	◯	◯	◯

——GOOD LUCK ——

SUPPLEMENTARY TABLE

Atomic No	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20
Symbol	H	He	Li	Be	B	C	N	O	F	Ne	Na	Mg	Al	Si	P	S	Cl	Ar	K	Ca
Mass No	1	4	7	9	11	12	14	16	19	20	23	24	27	28	31	32	35.5	40	39	40